

ATOMS AND IONS IN THE PERIODIC TABLE

AMPHIPHILIC ELEMENTS ARE INDICATED BY GREY SQUARES

IONIC RADIUS OF ELEMENTS BY Dr. J. A. Compton, Murray State College, Murray, Tennessee, Tennessee

INDICATED BY BLACK SQUARES

1. The radius of an atom is the distance from the nucleus to the outermost shell of electrons. It is measured in Angstroms (Å). The radius of an atom increases from left to right and from top to bottom in the periodic table.

2. The radius of an ion is the distance from the nucleus to the outermost shell of electrons. It is measured in Angstroms (Å). The radius of an ion increases from left to right and from top to bottom in the periodic table.

3. The radius of an atom is larger than the radius of an ion. This is because the ion has a smaller number of electrons than the atom, and the electrons are pulled closer to the nucleus.

4. The radius of an atom is larger than the radius of a cation. This is because the cation has a smaller number of electrons than the atom, and the electrons are pulled closer to the nucleus.

5. The radius of an atom is larger than the radius of an anion. This is because the anion has a larger number of electrons than the atom, and the electrons are pushed further away from the nucleus.

6. The radius of an atom is larger than the radius of a transition metal ion. This is because the transition metal ion has a smaller number of electrons than the atom, and the electrons are pulled closer to the nucleus.

7. The radius of an atom is larger than the radius of a noble gas atom. This is because the noble gas atom has a smaller number of electrons than the atom, and the electrons are pulled closer to the nucleus.

